

AP Chemistry Polyatomic Ions

| # | Polyatomic Ion and Oxidation number | Name |
|-----|---------------------------------------|-----------------------------------|
| 1. | ClO^{-1} | Hypochlorite |
| 2. | ClO_2^{-1} | Chlorite |
| 3. | ClO_3^{-1} | Chlorate |
| 4. | ClO_4^{-1} | Perchlorate |
| 5. | $\text{C}_2\text{H}_3\text{O}_2^{-1}$ | Acetate |
| 6. | OH^{-1} | Hydroxide |
| 7. | CN^{-1} | Cyanide |
| 8. | NO_2^{-1} | Nitrite |
| 9. | NO_3^{-1} | Nitrate |
| 10. | HCO_3^{-1} | Hydrogen carbonate or Bicarbonate |
| 11. | HSO_4^{-1} | Hydrogen sulfate or bisulfate |
| 12. | HSO_3^{-1} | Hydrogen sulfite or bisulfite |
| 13. | $\text{H}_2\text{PO}_4^{-1}$ | Dihydrogen phosphate |
| 14. | SCN^{-1} | Thiocyanate |
| 15. | IO_3^{-1} | Iodate |
| 16. | OCN^{-1} | Cyanate |
| 17. | MnO_4^{-1} | Permanganate |
| 18. | SO_4^{-2} | Sulfate |
| 19. | SO_3^{-2} | Sulfite |
| 20. | CO_3^{-2} | Carbonate |
| 21. | CrO_4^{-2} | Chromate |
| 22. | $\text{Cr}_2\text{O}_7^{-2}$ | Dichromate |
| 23. | SiO_3^{-2} | Silicate |
| 24. | $\text{C}_2\text{O}_4^{-2}$ | Oxalate |
| 25. | HPO_4^{-2} | Hydrogen Phosphate |
| 26. | O_2^{-2} | Peroxide |
| 27. | PO_4^{-3} | Phosphate |
| 28. | PO_3^{-3} | Phosphite |
| 28. | NH_4^{+1} | Ammonium |